

# Partial Pressure

When there is a mixture of gasses, each gas will exert a pressure in direct proportion to its representation in the mixture.



From Figure 2-1, page 14

Consider this hypothetical model:



Although the pressure in both containers is identical, 100 mm Hg, each gas will exert a partial pressure in direct proportion to its representation in the mixture. Each gas will move down its pressure gradient until an equilibrium is reached.

## **Atmospheric Pressures**

Atmospheric pressue at sea level is 760 mm Hg.

Gas	Partial Pressure	Percent Distribution In Mixture of Gases
PO <sub>2</sub> =	159 mm Hg	20.9%
PCO <sub>2</sub> =	0.3 mmHg	0.04 %
PN <sub>2</sub> =	597 mm Hg	78.6 %
$PH_2O =$	3.7 mm Hg	0.46 %
	760 mm Hg	100 %

#### Inter-Alveolar Pressures

Gas	Partial Pressure	Percent Distribution In Mixture of Gases
PO <sub>2</sub> =	104 mm Hg	13.7 %
$PCO_2 =$	40 mmHg	5.2 %
PN <sub>2</sub> =	569 mm Hg	74.9 %
$PH_2O =$	47 mm Hg	6.2 %
	760 mm Hg	100 %



Because ventilation does not result in a complete exchange of air due to significant "dead air space" in the conducting portion of the respiratory tree, gas proportions are significantly different between atmospheric air and inter-alveolar air.

Further, as the air is moisturized within the respiratory system, the percent representation from water vapor increases. The values reflect this increase.

Atmospheric Air

## External Respiration

(Between Alveolar Air and Blood)

When considering gas exchange,  $O_2$  and  $CO_2$  are of primary importance, therefore in this illustration and from here on out, other gasses will be ignored.



## **Internal Respiration**

(Between Blood and Interstitum)



Although the pressure in all chambers is the same, the gasses move in different directions down their pressure gradients.

